

## Synthesis and Microstructure Characterization of Novel Sr-HA Prepared by Co-precipitation with Enhanced Bioactivity

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**T**HERE IS accumulating evidence that strontium-containing biomaterials have positive effects on bone tissue repair. In this study, a series of Sr-substituted hydroxyapatites ( $\text{Sr}_x\text{Ca}_{1-x}\text{)}_5(\text{PO}_4)_3\text{OH}$ , where  $x = 0.00, 0.5, 1.00$  and  $2.00$  were made by the co-precipitation method. The microstructure analyses of obtained apatite after incorporation of Sr were evaluated by using X-ray Diffraction (XRD), Fourier transform infrared spectroscopy (FT-IR), Scanning electron microscope (SEM) and Thermo gravimetric analysis (TGA). Moreover, bioactivity of samples was examined by soaking samples in the simulated body fluid solution (SBF). Strontium is quantitatively incorporated into hydroxyapatite where its substitution for calcium provokes a linear increase in the lattice constants and a linear shift of the infrared absorption bands of the hydroxyl and phosphate groups, coherent with the greater ionic radius of strontium. Also incorporation of Sr into hydroxyapatite provokes its thermal stability and bioactivity. The obtained results indicate that Sr-HA have high potential to be used as a resorbable scaffold material release therapeutic agent to enhance osteogenesis and bone regeneration.

**Keywords:** Hydroxyapatite, Sr-substituted hydroxyapatite, Crystal size, Bioactivity, Simulated body fluid.

Worldwidly, the rise in life expectancy has led to an increased demand for new restoring and augmentation materials, with application in treatment therapies for bone defects<sup>(1)</sup>. These new compositions must tackle the bone regeneration therapies in a more efficient and durable way. The current scenario regarding treatments of bone disorders is that over 2.2 million persons are surgically treated each year as consequence of trauma or tumor extirpation<sup>(2)</sup>. The gold standard for grafting and repair is the reconstruction with autografts. However, the limitation of tissue amount and the morbidity at the extraction site have resulted in a high interest for synthetic materials<sup>(3)</sup>. An ideal bone graft should be reabsorbable and guide the patient's bone tissue towards a regenerative process<sup>(4)</sup>. Moreover, the graft should exhibit chemical, physical and biological properties appropriate to

the implantation site, especially when implanted as a solid piece or as porous scaffolds. Despite the research efforts in the field of ceramics, metals and polymers, the availability of this ideal material is still far away for clinical practice and a challenge<sup>(5)</sup>.

Bone is a dynamic and highly vascularized tissue that continues to remodel throughout the lifetime of an individual. It plays an integral role in locomotion, ensures the skeleton has adequate load-bearing capacity, and acts as a protective casing for the delicate internal organs of the body. In addition to these structural functions, bone is intimately involved in homeostasis through its storage of Ca and P ions and by regulating the concentration of key electrolytes in the blood<sup>(6)</sup>. Beside Ca and P, bone apatite is substituted by many different trace elements occurring in smaller concentration. Since many trace elements, such as Sr, Mg, Zn or Si present in the human body, are known for their anabolic effects in bone metabolism, new approaches for enhancing bioactivity of scaffold materials are being investigated by introducing therapeutic ions into the scaffold material<sup>(7)</sup>.

Strontium is known to accelerate bone healing processes and have positive effects on bone tissue repair. *In vitro* and *in vivo* studies have indicated that strontium increases bone formation and reduces osteoporosis, leading to a gain in bone mass and improved bone mechanical properties in normal animals and humans. Thus, research efforts are devoted to incorporate these Sr ions in synthetic apatites resulting in the changed dissolution behaviour of these materials and their modified (usually improved) biological performance<sup>(8)</sup>.

In this study, Sr substituted hydroxyapatite ceramics were prepared by the co-precipitation method. The effect of incorporation of Sr ions on both physical and biological properties of hydroxyapatite was reported. The prepared samples were characterized using different techniques. Moreover, the bioactivity and biodegradability tests in a simulated body fluid (SBF) were evaluated.

## Materials and Methods

### *Preparation of Sr-HA*

Non-substituted and Sr-substituted hydroxyapatites are prepared by using the aqueous precipitation method. Pure hydroxyapatite (HA) was prepared by the aqueous precipitation method from Calcium nitrate tetra hydrate and ammonium dihydrogen phosphate as the starting material. Calculated amounts of  $(\text{Ca}(\text{NO}_3)_2 \cdot 4\text{H}_2\text{O})$  and  $((\text{NH}_4)\text{H}_2\text{PO}_4)$  are dissolved in 1000 ml of distilled water by taking into consideration that Ca/P molar ratio is adjusted to about 1.67 and pH for each aqueous solution is kept at 10 -11 by the addition of ammonia solution when it is required. All the dissolution is performed at room temperature. After dissolution, the temperature of vigorously stirred  $(\text{Ca}(\text{NO}_3)_2 \cdot 4\text{H}_2\text{O})$  solution is raised until it reached 60 °C, after that, the vigorously stirred  $((\text{NH}_4)\text{H}_2\text{PO}_4)$  solution is added drop wise to the  $(\text{Ca}(\text{NO}_3)_2 \cdot 4\text{H}_2\text{O})$  solution at a pH kept at 10-11. When dropping is completed, the temperature of solution is raised to about 100 °C. After it reached 100°C, the mother solution is kept for about 2 hours on the

stirrer. After that, the solution is aged overnight on which the precipitation of the solution is formed. On the next day the precipitated solution was filtered and washed several times with distilled water. Then it was dried at about 110°C and sintered at about 600 °C. The sintered powders are sieved at 150  $\mu\text{m}$ <sup>(9)</sup>. Strontium substituted hydroxyapatite was prepared by an analogous method, but with a reduction of the amount of  $\text{Ca}^{+2}$  and an addition of a calculated amount of  $\text{Sr}^{+2}$  in the form of strontium nitrate (( $\text{Sr}(\text{NO}_3)_2$  Sigma). Four different compositions were prepared which are named HA (pure), **S<sub>1</sub>** (0.5 wt % of Sr), **S<sub>2</sub>** (1 wt % of Sr) and **S<sub>3</sub>** (2 wt % of Sr).

#### Characterization techniques

##### X-ray diffraction (XRD)

Phase analysis of the prepared samples are studied using X-ray diffractometer (XRD, Bruker D8 Advanced Cu target with secondary monochromator) operated at 40 kV and tube current of 40 mA. Diffraction patterns were compared to ICDD database PDF patterns of  $\text{Ca}_{10}(\text{PO}_4)_6(\text{OH})_2$  (HA), (JCPDS No.024-0033).

The crystal size of the precipitates was estimated from the XRD pattern using the Scherrer equation.

$$D_{hkl} = \frac{K\lambda}{\beta \cos\theta_{hkl}} \rightarrow \text{-----} 1$$

where K is a constant varying with the method of measuring the broadening and is chosen to be 0.9,  $\lambda$  is the wavelength (nm) of Cu  $K_\alpha$  radiation ( $\lambda = 0.15418$  nm),  $\beta$  corresponds to full width at half maximum (FWHM) for the peak hkl (rad) and  $\theta$  is the diffraction angle (in degrees)<sup>(10)</sup>.

The degree of crystallinity corresponding to the fraction of crystalline phase present in the examined volume can be estimated by X-ray diffraction data according to the following:

$$X_c = 1 - \frac{V_{112/300}}{I_{300}} \rightarrow \text{-----} 2$$

where  $X_c$  is the crystallinity degree,  $I_{300}$  is the intensity of (300) reflection and  $V_{112/300}$  is the intensity of the hollow between (112) and (300) reflection<sup>(11,12)</sup>.

##### Fourier transform infrared spectroscopy (FT-IR)

Functional groups present on the prepared samples can be determined using the (FT-IR) technique. The (FT-IR) spectra are obtained in the region of 4000-400  $\text{cm}^{-1}$  (wave number) using KBr pellet transmission technique with a resolution of 4  $\text{cm}^{-1}$  where 0.002 g of powder samples were mixed with 0.198 g of KBr and then pressed into pellets.

#### *Thermal gravimetric analysis (TGA)*

The weight changes of the prepared (uncalcined) powders were studied as a function of temperature by the thermal gravimetric analysis (TGA), Shimadzu TGA-50 Hz instrument. The test was performed at a heating rate of 10°C/min under nitrogen atmosphere from RT up to 1000 °C.

#### *Apatite formation ability*

The prepared powders are processed in the form of discs (HA and S<sub>3</sub>) and then immersed in SBF solution that was prepared according to Kokubo and Takadama<sup>(13)</sup>. The samples are soaked for different periods; 1, 2, 6 and 12 days, at 37°C. The solutions are analyzed using the Inductively coupled plasma (ICP) that measure the change in Ca<sup>+2</sup>, PO<sub>4</sub><sup>-3</sup> and Sr<sup>+2</sup> ions concentration during soaking the samples in SBF (ICP, Model Ultima 2-Jobin Yvon) which operates between 1 and 5 kilowatts. Moreover, the pH of the SBF is measured using a pH meter (Fisher Scientific, Pittsburgh PA).

#### *Scanning electron microscope (SEM)*

The surface of prepared samples before and after soaking in SBF are investigated by scanning electron microscope (SEM, Jeol JxA 840). The samples were sputtered with gold by a sputter coater as an adhesive electric conductor.

## **Results and Discussion**

#### *Phase identification by using XRD*

The XRD analysis of the HA and Sr-HA powder (Fig. 1), revealed that, they have a hexagonal shape showing the characteristic peaks of hydroxyapatite structure without secondary phase formation to detect due to the presence of Sr which demonstrates the complete substitution of Sr in the hydroxyapatite network.

The average crystallite size of the samples is calculated using the Scherrer formula (No.1.), as shown in Table1. At low Sr content (0.5 wt %, 1 wt %), crystal size of Sr-HA decreased as compared with that of HA. This may be due to at low strontium contents, Sr prefer to present only in a limited number of unit cells, and it is present in only a fraction of M(1) site as this site allows the accommodation of a larger cation because of the longer bonds M(1)–O mean distances. However, when the number of heavier ions increases (2wt% Sr content), the repulsion between atoms in the M(1) position increases that causes Sr to prefer to occupy M(2) sites as the strontium replacement of the calcium at M(2) sites allows a better accommodation of the heavier strontium atoms because in this position metal atoms form “staggered” equilateral triangles centered on the unit cell of apatite channel<sup>(14)</sup>.

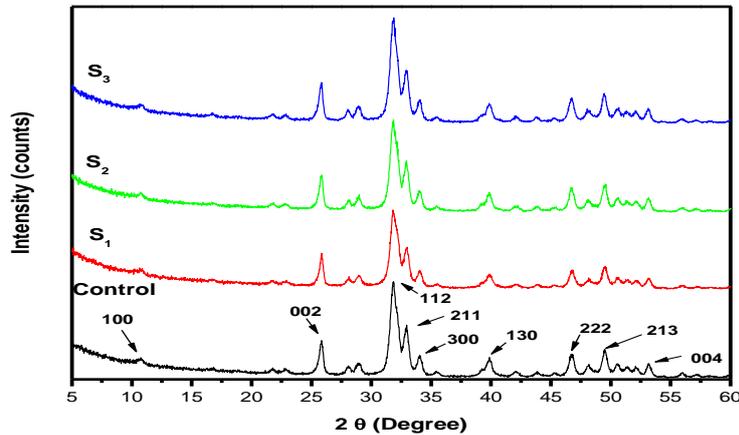


Fig.1. XRD of control (HA) and Sr-HA sintered at about 600 °C.

TABLE 1. Crystal size, crystallinity and lattice parameters of Sr-HA reflected by XRD patterns

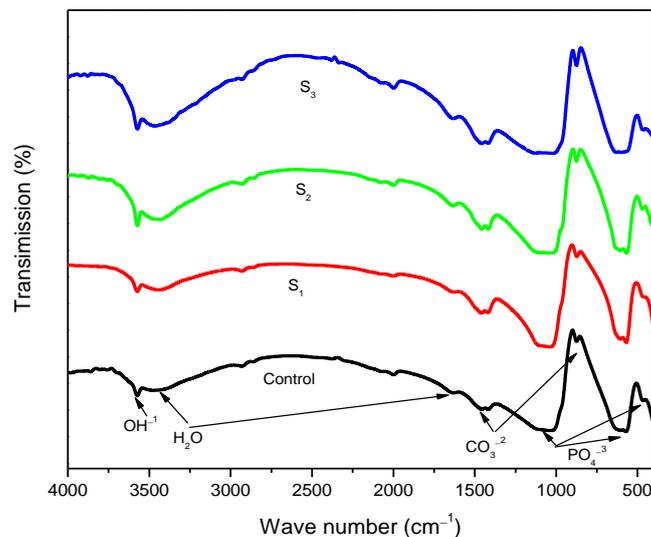
Sample	Sr content (%)	Crystallinity $x_c$ (%)	Average Crystal dimension $d$ (nm)	Lattice parameters	
				a	c
control	0	52	29.12	9.11	6.9014
S <sub>1</sub>	0.5	47.58	25.30	9.12	6.9016
S <sub>2</sub>	1	52.6	25.12	9.17	6.9022
S <sub>3</sub>	2	53	29.59	9.18	6.9034

The degree of crystallinity (calculated by No. 2) increases from 52% for control to 52.6-53% for S<sub>2</sub> and S<sub>3</sub>, respectively but decreased to 47.58 for S<sub>1</sub> which matched with the observed broadening in the XRD peaks at 0.5% Sr content. The broadening of the X-ray line width indicated that the incorporation of Sr at this content destroyed the symmetry. This may be due to a greater difficulty for HA to host the larger strontium ion than for SrHA to host the smaller calcium ion (Ca<sup>+2</sup> ionic radius = 0.100 nm; Sr<sup>+2</sup> ionic radius = 0.118 nm)<sup>(14)</sup>.

The effect of ion substitution on the crystal structure of hydroxyapatite is determined by detecting a variation in the lattice parameters (calculated by No. 3). The ionic radius of Sr is larger than that of Ca which is in agreement with the observed increase in lattice parameters for Sr-HA as shown in Table 1. The change of lattice parameters of Sr-HA clearly demonstrated that Sr ion was structurally incorporated, in other words, they didn't just cover the surface of crystal<sup>(15)</sup>.

*FT-IR spectroscopic studies*

Figure 2 shows the FTIR spectra of pure hydroxyapatite and Sr-HA samples. All samples exhibit the characteristic pattern of partially carbonated hydrated hydroxyapatite, as extensively discussed elsewhere. The absorption bands at 1040, 570 and 466  $\text{cm}^{-1}$ , detected in the spectrum, are attributed to the phosphate ( $\text{PO}_4^{3-}$ ) characteristic absorption, were presented in all spectra for the synthetic apatites<sup>(16)</sup>. The  $\text{OH}^-$  stretching mode is observed at 3573  $\text{cm}^{-1}$ , this band in the spectrum of  $S_1$  has a lower intensity than those in  $S_2$  and  $S_3$  while  $\text{OH}^-$  vibration mode is disappeared in all prepared samples<sup>(17)</sup>. The FTIR analysis shows the presence of B-type carbonation ( $\nu_3$  carbonate vibrational mode) in the range 1417-1457  $\text{cm}^{-1}$ . Although there is no carbonate source in the solution such presence of carbonation on the prepared samples may be due to the carbon dioxide in the reaction vessel, since the synthesis process was performed under normal atmosphere<sup>(15)</sup>. It has been reported previously that co-substitution of Sr and carbonate in a calcium phosphate lattice could enhance the biocompatibility properties and dissolution rate of hydroxyapatite<sup>(18)</sup>. The band at about 875  $\text{cm}^{-1}$  may be attributed to  $\text{HPO}_4^{2-}$ . The band due to adsorbed water was noticed at 1632  $\text{cm}^{-1}$  while the band that observed at 3479  $\text{cm}^{-1}$ . This observed band of adsorbed water may be referred to its entrance into the lattice of HA.



**Fig .2.** Infrared spectra of the Sr-HA prepared samples sintered at about 600 °C.

Figure 2 shows how the bands due to the symmetric stretching ( $\nu_3$ ) and bending ( $\nu_4$ ) modes of phosphate groups shifts to lower wave number .The predominant factor causing shifts in the internal phosphate frequencies to lower energies is decreased anion–anion repulsion concomitant with an increased anion–anion separation on

increasing cation radius. Such changes may be attributed to the substitution of  $\text{Sr}^{+2}$  for  $\text{Ca}^{+2}$  into the lattice of apatite, the greater ionic radius of  $\text{Sr}^{+2}$  consequently decreases the bonding strength of P-O, the vibrations further confirm that  $\text{Sr}^{+2}$  can substitute  $\text{Ca}^{+2}$  and enter the lattice of apatite<sup>(19)</sup>. Similar observations were also made by other authors in the literature<sup>(17,20)</sup>. Also the intensity of  $\nu_2$  phosphate band and of  $\text{OH}^{-1}$  stretching mode is increased with increasing Sr content but decreased on the spectrum of  $\text{S}_1$ . This may be due to the increase on the degree of crystallinity with increasing Sr content for high Sr ratio (1%, 2%) but a decrease on crystallinity for low Sr ratio (0.5%). This results have been matched with the results obtained from the XRD analysis for Sr-HA.

#### Thermogravimetric analysis (TGA)

The TGA plots illustrated in Fig.3 describe the weight loss along the investigated temperature range for the prepared samples. From this figure, it is observed that the mass loss during heating could be classified into two stages. The first stage weight loss occurs at 27 to 370°C due to the vaporization of physisorbed and chemisorbed water from the apatite lattice that's nearly about 7%<sup>(21,22)</sup>. Water monolayer, that is in contact with the HA surface (chemisorbed layer), was more strongly bound than the additional water layers (all physisorbed layers) that involved water/water contacts only. Chemisorbed water are found to be completely removed from the HA at about 300°C, whereas physisorbed water are removed at about 20°C<sup>(23)</sup>. The second stage weight loss occurs at ~ 370 to 1000°C. The mass fraction lost in this temperature range was about 3wt%. As shown from the figure, temperature of phase transformation of HA  $\text{S}_1 < \text{S}_2 < \text{S}_3$  which means that, Sr incorporation into HA increases its thermal stability.

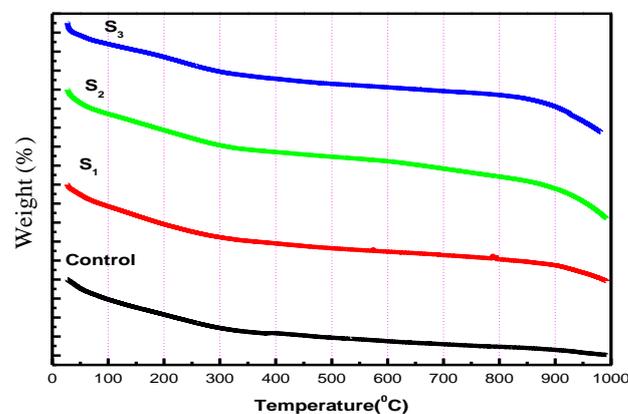
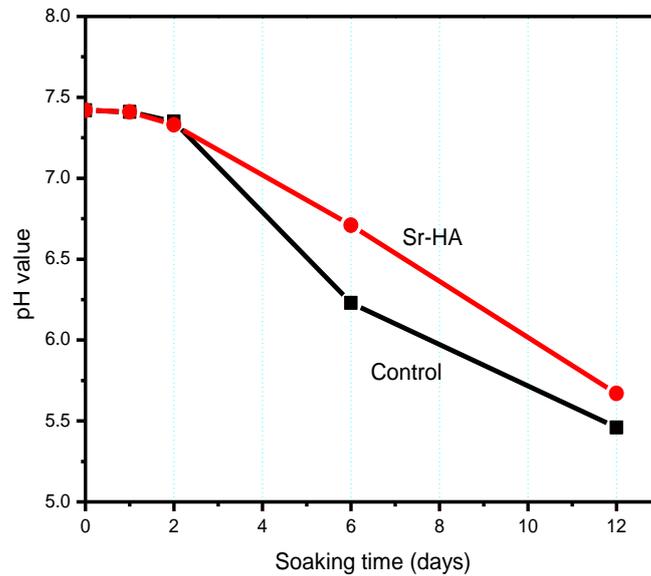


Fig.3. TGA curves of control (HA) and Sr-HA samples as dried .

#### Bioactivity test (Activity of prepared samples in SBF)

Results of pH measurements for disc samples immersed for different periods in SBF solution are shown in Fig. 4. All discs of the different hydroxyapatites showed a decrease in the value of pH<sup>(23)</sup>.



**Fig. 4. Changes in SBF pH with time of immersion for control (HA) and Sr-HA samples (S<sub>3</sub>).**

Figure 5.a and 5.b show calcium ( $\text{Ca}^{+2}$ ), and phosphorus ( $\text{P}^{+5}$ ) ion concentrations in the SBF solution after the different immersion periods, of synthesized apatite samples. As shown in those figures,  $\text{Ca}^{+2}$  and  $\text{P}^{+5}$  ions concentration decreased in the early stages for the prepared powders. This diminution is mainly due to uptake of these ions from SBF by the prepared samples to use them in the formation of carbonated (bone like) apatite layer on their surfaces. This decrement is followed by an enlargement in the ions concentration which means that the release of these ions from the samples into SBF solution is over than the uptake of calcium from SBF (degradation of samples overcomes the formation of the apatite layer)<sup>(13)</sup>.

Figure 5.c. shows  $\text{Sr}^{+2}$  ion concentration through the periods of immersion. As shown from the figure, there was an increase in the Sr ion concentration for the first period of immersion; this may be due to the release of Sr from the Sr-HA samples in which the SBF solution does not contain any Sr content. Starting to increase is then followed by a decrease in the Sr ion concentration due to consumption of Sr ions from the SBF solution and their incorporation for the formation of an apatite layer.

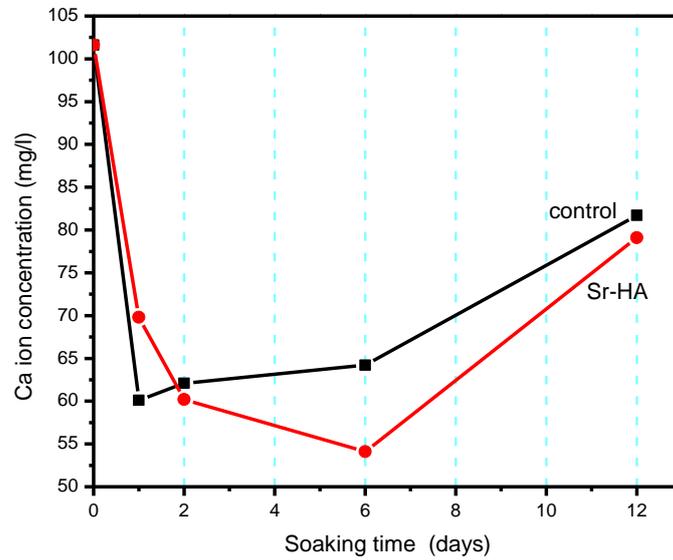


Fig. 5a. Calcium ion concentration in SBF of control (HA) and Sr-HA .

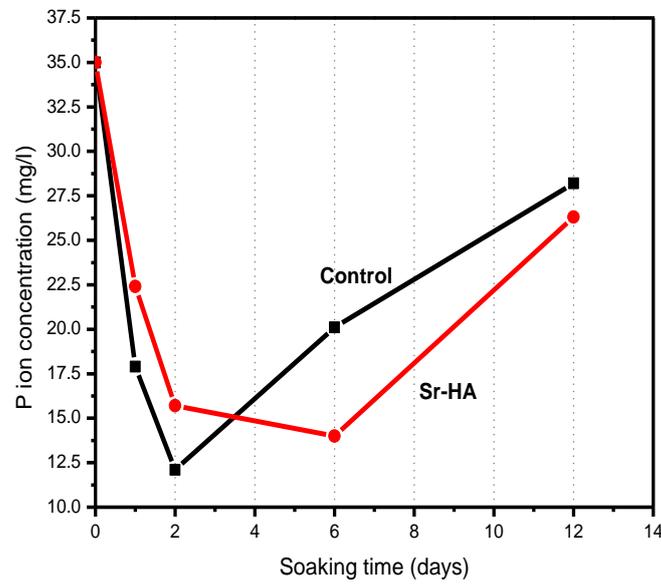
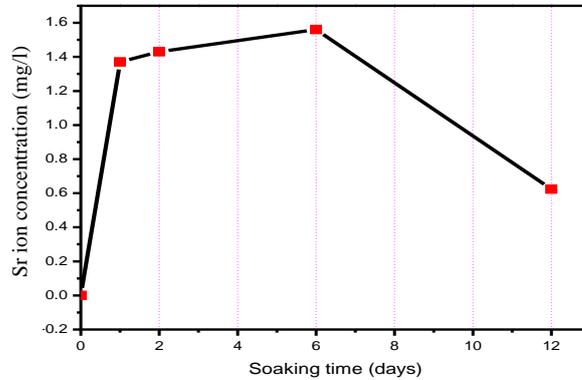
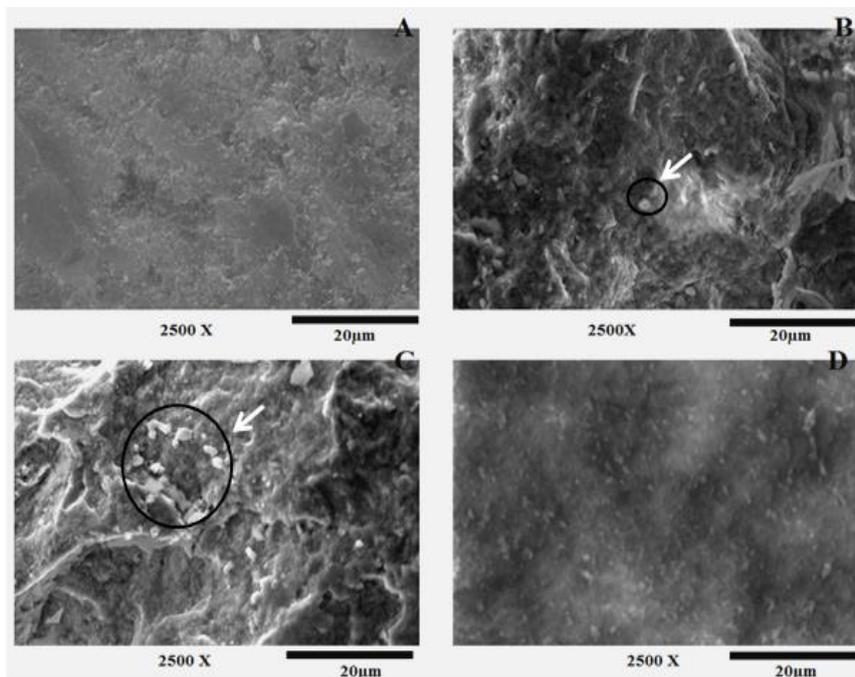


Fig. 5b. Phosphorus ion concentration in SBF of control (HA) and Sr-HA .



**Fig 5.c. Strontium ion concentration in SBF of Sr-HA.**

SEM of processed disc samples before and after the immersion in SBF for 1, 2 and 12 days are shown in Fig. 6 A new layer of nano-sized precipitates agglomerated in clusters is formed. They appear as bright tiny spots. As the soaking time was increased, the number and size of the agglomerated particles also increased. The increase of agglomerated particles is evident due to the formation of apatite or mineralization being taking place on the surface of the prepared Sr-HA samples.



**Fig.6. SEM of Sr-HA .A) Before immersion in SBF, B) After immersion for 1 day, C) After immersion for 2 day, D) after immersion for 12 day.**

### Conclusion

Sr-HA solid solutions with various Sr content were prepared by the co-precipitation method. Increasing strontium substitution for calcium in the HA structure provokes a linear variation in the cell parameters and in the infrared absorption bands, in agreement with the increasing mean dimensions of the cation. At variance, the effect of strontium on crystallinity and morphology changes with composition: Relatively low Sr replacement to calcium induces a decrease of the coherent length of the perfect crystalline domains and disturbs the shape of the crystals, whereas crystallinity, as well as mean dimensions of the crystals, increases at relatively high strontium contents. The incorporation of Sr into the lattice of HA evidently increases its thermal stability and bioactivity. Thus, this study shows that the development of strontium containing apatite represent a promising bone-grafting material for bone regeneration procedures.

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### تحضير وتوصيف عينات من الهيدروكسي أباتيت المبدل بعنصر سترنشيوم والتميز بالنشاط الحيوى بطريقة الترسيب المشترك

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يعتبر عنصر الاسترنشيوم من أهم العناصر التى يمكن أن يبدل بها عنصر الكالسيوم فى مركب الهيدروكسي أباتيت وقد حظى هذا العنصر باهتمام كبير نظرا لأهميته فى عملية تكوين العظام حيث أنه يعمل على زيادة معدل البناء ويقلل من عملية الهدم. تم فى هذه الدراسة فحص تأثير إبدال عنصر الكالسيوم بعنصر الاسترنشيوم على خواص مركب الهيدروكسي أباتيت حيث حل الاسترنشيوم محل ٥,٠% , ١% و ٢% بالوزن من نسبة الكالسيوم وقد تم تحضير العينات بطريقة الترسيب المشترك. ثم تم دراسة التغيرات التى طرأت على الخواص الكيميائية والتركيب البلورى لمركب الهيدروكسي أباتيت بعد إدخال عنصر الاسترنشيوم به لتوضيح التغيرات التى قد تطرأ على تركيب العظام عندما يتواجد به عنصر الاسترنشيوم. وقد تم تحليل النتائج باستخدام ظاهرة حيود أشعة إكس لتوضيح التغيرات التى تطرأ على التركيب البلورى والتحول الطورى لمركب الهيدروكسي أباتيت بعد تضمينه لعنصر الاسترنشيوم. أما طريقة التحليل الطيفى بالأشعة تحت الحمراء فقد استخدمت لبيان التغيرات فى مجموعات المركبات الكيميائية لمركب الهيدروكسي أباتيت. تم دراسة الخواص الحرارية للمركب باستخدام طريقة التحليل الحرارى الوزنى. وقد تم تحضير سائل محاكى لسائل جسم الانسان لدراسة النشاط الحيوى للعينات تحت البحث. وقد خلص هذا البحث إلى أن وجود عنصر سترنشيوم قد عمل على تحسين خواص مركب الهيدروكسي أباتيت مما يتيح استخدامه فى إنتاج دعامات حيوية تستخدم فى التطبيقات الطبية.